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CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C 3H 4(g) + xO 2(g) → 3CO 2(g) + 2H 2O(g) 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C 3H 4? 2 mol O 2:1 mol H 2O c. What is the mole ratio of O 2 to H

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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H

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Stoichiometry b. Theoretically, how many moles of NH3 will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N2 are mixed with 12.0 mol of H2 according to the ...

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry ... 290 Chapter 9 DO NOT EDIT--Changes must be made through “File info” ... The number of significant figures in the answer

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Play this game to review Chemistry. Avogadro’s number is: Q. Using the pictured equation, how many grams of zinc chloride are produced from 7.89 moles of zinc?

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Chapter 9 - Stoichiometry. 9-1 Introduction to Stoichiometry. Composition Stoichiometry - deals with mass relationships of elements in compounds Reaction Stoichiometry - Involves mass relationships between reactants and products in a chemical reaction. I. Reaction Stoichiometry Problems A. Four problem Types, One Common Solution.

Chapter 9 - Stoichiometry

Ch. 9 Review: Stoichiometry KEY Page 1 1. The following equation represents a laboratory preparation for oxygen gas: 2KClO3(s) + heat 2KCl(s) + 3O2(g) How many moles of O2 form as 3.0 mol of KClO3 are totally consumed? 3.0 mol KClO3 x (3 moles O2)/(2 moles KClO3) = 4.5 moles O2 2.

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Chapter 9 - Stoichiometry Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a...