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The separation of

solute particles (ΔH_1

Solvent-solvent

Interactions...

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the water m...

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and Weaknesses of
Models □ The Valence-
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One Possible Structure
12.3 Resonance

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chemistry. law of
conservation of mass.
formula mass. law of
definite proportions.
limiting reactant. no
detectable change in

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mass occurs during a chemical reaction. the sum of the atomic masses given in the formula of a substan....

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Review Skills 12.1

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Liquid and from Liquid
to Gas - An

Introduction to
Dynamic Equilibrium

The Process of
Condensation The
Process of Evaporation

How Evaporation
Causes Cooling Rate of
Evaporation Dynamic

Equilibrium Between
Liquid and Vapor

Equilibrium Vapor

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Evaporation, and
Dynamic ...**

Section 12.1 Gases 1.
motion 2. a. small b.
forces c. random d.
elastic; kinetic 3. KE
 $\frac{1}{2}mv^2$ 4.
Temperature 5. true 6.
true 7. false 8. true 9.
true 10. false 11. true
12. false 13. a 14. a 15.
d 16. d 17. b 18. b 19.
b 20. barometer 21.

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Evangelista Torricelli

22. gravity and

atmospheric pressure

23. The atmospheric

pressure has

increased. Section 12

.2 Forces of Attraction

1. intermolecular

forces 2.
intermolecular forces

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CREATOR. Religion
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Terms.

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PLAY. Match. Gravity.
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this set (33) Carbon.
Organic chemistry is
the study of all
substances containing.
Inorganic chemistry.

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The branch of chemistry that deals with products without carbon is called.

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Matter Section 12.1

Gases 1. motion 2. a.

small b. forces c.

random d. elastic;

kinetic 3.

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Alcohols Alcohols are
organic compounds
that have a hydroxyl
(-OH)

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calculation of
quantities in chemical
reactions. The

quantities to consider
in a chemical reaction?

On the left side of a
chemical equation, you
have the? On the right
side of a chemical

equation, you have
the? In every chemical
reaction, these two
things are conserved.

In a majority of
chemical reactions,

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these three things are NOT conserved.

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